

QUIZIZZ

Electron Configurations

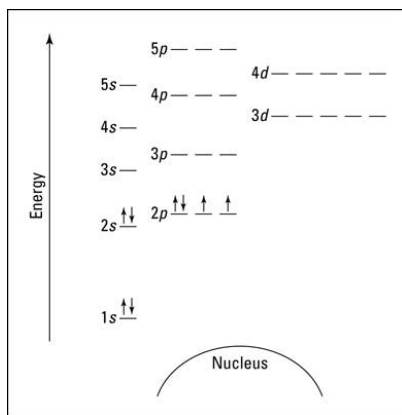
16 Questions

NAME : _____

CLASS : _____

DATE : _____

1.



How many electrons can the first energy level hold?

a) 1

b) 2

c) 8

d) 0

2. The electron configuration of an atom is $1s^2 2s^2 2p^6$. The number of electrons in the atom is

a) 3

b) 6

c) 8

d) 10

3.

Which electron configuration belongs to Chlorine (Cl)?

a) $1s^2 2s^2 2p^6 3s^2 3p^5$

b) $1s^2 2s^2 2p^6 3s^2 3p^6$

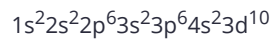
c) $1s^2 2s^2 2p^6 3s^2 3p^7$

Help

4.

Periodic Table of the Elements

What atom matches this electron configuration?



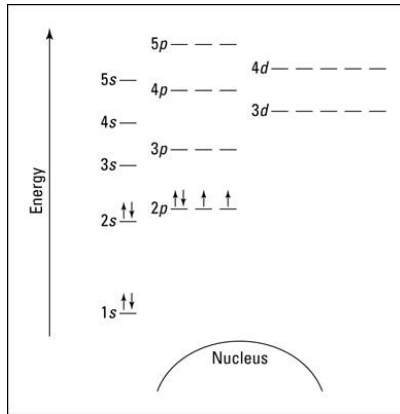
a) Zinc

b) Copper

c) Nickel

d) Germanium

5.



What electron configuration matches an oxygen atom?

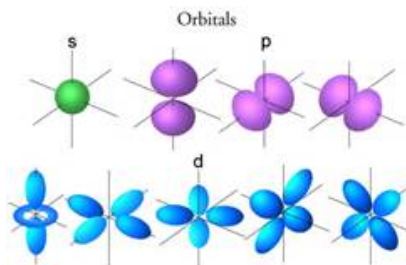
a) $1s^2 2s^2 2p^6 3s^2, 3p^6 4s^2 3d^{10} 4p^5$

b) $1s^2 2s^2 2p^4$

c) $1s^2 2s^2 2p^6$

d) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^1$

6.



How many electrons can the d sublevel hold?

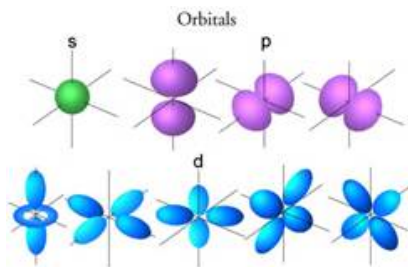
a) 14

b) 10

c) 2

d) 6

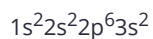
7.



How many electrons can the s sublevel hold?

 a) 14 b) 10 c) 2 d) 6

8. What atom matches this electron configuration?

 a) Neon b) Magnesium c) Aluminum d) Potassium

9. Which area on the periodic table represents the electrons in the "d" sublevel?

 a) representative elements b) columns 3-8 c) rare earth elements d) transition elements

10. Each row on the periodic table represents:

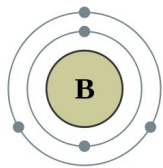
 a) an energy level b) a sublevel c) an electron d) an orbital

11.

5: Boron

2,3

What's a valence electron?

 a) electrons in the second energy level b) electrons in the outermost energy level c) the atomic number d) electrons in the first level

12. How many valence electrons are in Phosphorus?

 a) 15 b) 4 c) 5 d) 31

13.



How many electrons are in the outer shell of Sodium (Na)?

 a) 3 b) 1 c) 2 d) 4

14. How many electrons should Beryllium have around its Lewis dot model?

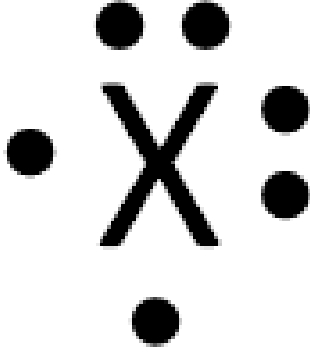
 a) 1 b) 2 c) 3 d) 4

15. How many electrons should Fluorine have around its Lewis dot model?

 a) 5 b) 6 c) 7 d) 8

16.

This could be the dot diagram of



a) Mg

c) C

b) Cl

d) O

Answer Key

1. b
2. d
3. a
4. a

5. b
6. b
7. c
8. b

9. d
10. a
11. b
12. c

13. b
14. b
15. c
16. d