

Assignment 6 Ionic Binary Compounds (d block + nonmetal) KEY

Write the name of each of the following compounds.

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|------------------------------------|--------------------------|
| 1. CuF | copper (I) fluoride |
| 2. CuF ₂ | copper (II) fluoride |
| 3. Cr ₂ O ₃ | chromium (III) oxide |
| 4. PbI ₂ | lead (II) iodide |
| 5. ZnCl ₂ | zinc chloride |
| 6. CrO ₃ | chromium (VI) oxide |
| 7. AuBr | gold (I) bromide |
| 8. NiO | nickel (II) oxide |
| 9. VI ₃ | vanadium (III) iodide |
| 10. SnO ₂ | tin (IV) oxide |
| 11. Mn ₂ O ₇ | manganese (VII) oxide |
| 12. NbCl ₅ | niobium (V) chloride |
| 13. Ti P | titanium (III) phosphide |
| 14. PaS ₂ | palladium (IV) sulfide |
| 15. PtF ₂ | platinum (II) fluoride |
| 16. Os ₂ O ₃ | osmium (III) oxide |
| 17. Ir ₃ N ₄ | iridium (IV) nitride |
| 18. CoCl ₂ | cobalt (II) chloride |
| 19. Fe ₂ S ₃ | Iron (III) sulfide |
| 20. AuI ₃ | gold (III) iodide |

<u>Ions</u>	Chloride Cl ¹⁻	Fluoride F ¹⁻	Nitride N ³⁻	Sulfide S ²⁻	Selenide Se ²⁻	Bromide Br ¹⁻	Phosphide P ³⁻
Gold (I) Au ¹⁺	AuCl	AuF	Au ₃ N	Au ₂ S	Au ₂ Se	AuBr	Au ₃ P
Gold (II) Au ²⁺	AuCl ₂	AuF ₂	Au ₃ N ₂	AuS	AuSe	AuBr ₂	Au ₃ P ₂
Manganese (II) Mn ²⁺	MnCl ₂	MnF ₂	Mn ₃ N ₂	MnS	MnSe	MnBr ₂	Mn ₃ P ₂
Manganese (IV) Mn ⁴⁺	MnCl ₄	MnF ₄	Mn ₃ N ₄	MnS ₂	MnSe ₂	MnBr ₄	Mn ₃ P ₄
Manganese (VII) Mn ⁷⁺	MnCl ₇	MnF ₇	Mn ₃ N ₇	Mn ₂ S ₇	Mn ₂ Se ₇	MnBr ₇	Mn ₃ P ₇
Chromium (II) Cr ²⁺	CrCl ₂	CrF ₂	Cr ₃ N ₂	CrS	CrSe	CrBr ₂	Cr ₃ P ₂
Chromium (III) Cr ³⁺	CrCl ₃	CrF ₃	CrN	Cr ₂ S ₃	Cr ₂ Se ₃	CrBr ₃	CrP
Tin (II) Sn ²⁺	SnCl ₂	SnF ₂	Sn ₃ N ₂	SnS	SnSe	SnBr ₂	Sn ₃ P ₂
Tin (IV) Sn ⁴⁺	SnCl ₄	SnF ₄	Sn ₃ N ₄	SnS ₂	SnSe ₂	SnBr ₄	Sn ₃ P ₄
Lead (II) Pb ²⁺	PbCl ₂	PbF ₂	Pb ₃ N ₂	PbS	PbSe	PbBr ₂	Pb ₃ P ₂
Lead (IV) Pb ⁴⁺	PbCl ₄	PbF ₄	Pb ₃ N ₄	PbS ₂	PbSe ₂	PbBr ₄	Pb ₃ P ₄
Iron (II) Fe ²⁺	FeCl ₂	FeF ₂	Fe ₃ N ₂	FeS	FeSe	FeBr ₂	Fe ₃ P ₂
Iron (III) Fe ³⁺	FeCl ₃	FeF ₃	FeN	Fe ₂ S ₃	Fe ₂ Se ₃	FeBr ₃	FeP
Copper (I) Cu ¹⁺	CuCl	CuF	Cu ₃ N	Cu ₂ S	Cu ₂ Se	CuBr	Cu ₃ P
Copper (II) Cu ²⁺	CuCl ₂	CuF ₂	Cu ₃ N ₂	CuS	CuSe	CuBr ₂	Cu ₃ P ₂