

Mole Practice Problem Set #2

NAME _____

You must show work on a separate piece of paper in order to get credit.

1. 59 g of Hydrochloric Acid = _____ moles
2. 0.46 moles of Sodium Hydroxide = _____ g
3. 35g of Sulfuric Acid = _____ mol
4. 877 g of Potassium Permanganate = _____ mol
5. 3.4×10^{24} moles of Copper (II) Bromide = _____ g
6. 0.56 moles of Calcium Hydroxide = _____ g
7. 4.0×10^{-23} g of Magnesium = _____ mol
8. 49 g of Dihydrogen Monoxide = _____ mol
9. 6.4 g of Iron (III) Chlorate = _____ mol
10. 4.5 moles of Nitric Acid = _____ g
11. 2.1×10^{-10} g of Phosphorus Pentabromide = _____ mol
12. 9.6 moles of Nickel (II) Sulfate = _____ g
13. 1.15 moles of Acetic Acid = _____ g
14. 65.7 g of Silver Nitrate = _____ mol
15. 0.25 moles of Strontium Phosphide = _____ g