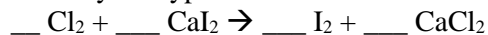


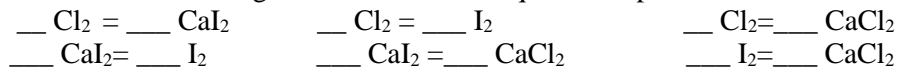
Chemical Reactions**STOICHIOMETRY INTRO WORKSHEET**

1. Name the following substances:			
	KClO ₃ :	KCl:	O ₂ :
2. Find the molar mass of each substance:			
	KClO ₃ :	KCl:	O ₂ :
3. Given 0.58 mol sample of O ₂ (g) at STP.			
(a)	Find the mass of this sample.		
(b)	Find the volume of this sample.		
(c)	Find the number of molecules in this sample.		
4. Identify the type of reaction and balance the equation. $\underline{\hspace{1cm}} \text{KClO}_3 \rightarrow \underline{\hspace{1cm}} \text{KCl} + \underline{\hspace{1cm}} \text{O}_2$			
5. Give the following mole ratios for the equation in problem #4 $\underline{\hspace{1cm}} \text{KClO}_3 = \underline{\hspace{1cm}} \text{KCl} \qquad \underline{\hspace{1cm}} \text{KCl} = \underline{\hspace{1cm}} \text{O}_2 \qquad \underline{\hspace{1cm}} \text{KClO}_3 = \underline{\hspace{1cm}} \text{O}_2$			
6. Identify the type of reaction and balance the equation. $\underline{\hspace{1cm}} \text{K} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{K}_2\text{O}$			
7. Give the following mole ratios for the equation in problem #6 $\underline{\hspace{1cm}} \text{K} = \underline{\hspace{1cm}} \text{O}_2 \qquad \underline{\hspace{1cm}} \text{K} = \underline{\hspace{1cm}} \text{K}_2\text{O} \qquad \underline{\hspace{1cm}} \text{O}_2 = \underline{\hspace{1cm}} \text{K}_2\text{O}$			
8. Identify the type of reaction and balance the equation. $\underline{\hspace{1cm}} \text{CH}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$			
9. Give the following mole ratios for the equation in problem #8 $\underline{\hspace{1cm}} \text{CH}_4 = \underline{\hspace{1cm}} \text{O}_2 \qquad \underline{\hspace{1cm}} \text{CH}_4 = \underline{\hspace{1cm}} \text{H}_2\text{O} \qquad \underline{\hspace{1cm}} \text{O}_2 = \underline{\hspace{1cm}} \text{CO}_2$ $\underline{\hspace{1cm}} \text{CO}_2 = \underline{\hspace{1cm}} \text{H}_2\text{O} \qquad \underline{\hspace{1cm}} \text{CH}_4 = \underline{\hspace{1cm}} \text{CO}_2 \qquad \underline{\hspace{1cm}} \text{O}_2 = \underline{\hspace{1cm}} \text{H}_2\text{O}$			
10. Identify the type of reaction and balance the equation. $\underline{\hspace{1cm}} \text{Li} + \underline{\hspace{1cm}} \text{Mg}_3\text{N}_2 \rightarrow \underline{\hspace{1cm}} \text{Li}_3\text{N} + \underline{\hspace{1cm}} \text{Mg}$			
11. Give the following mole ratios for the equation in problem #10 $\underline{\hspace{1cm}} \text{Li} = \underline{\hspace{1cm}} \text{Mg}_3\text{N}_2 \qquad \underline{\hspace{1cm}} \text{Li} = \underline{\hspace{1cm}} \text{Li}_3\text{N} \qquad \underline{\hspace{1cm}} \text{Li} = \underline{\hspace{1cm}} \text{Mg}$ $\underline{\hspace{1cm}} \text{Mg}_3\text{N}_2 = \underline{\hspace{1cm}} \text{Li}_3\text{N} \qquad \underline{\hspace{1cm}} \text{Mg}_3\text{N}_2 = \underline{\hspace{1cm}} \text{Mg} \qquad \underline{\hspace{1cm}} \text{Li}_3\text{N} = \underline{\hspace{1cm}} \text{Mg}$			

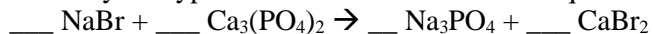
12. Identify the type of reaction and balance the equation.



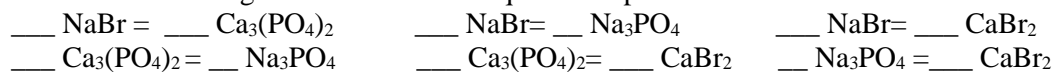
13. Give the following mole ratios for the equation in problem #12



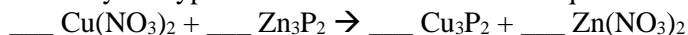
14. Identify the type of reaction and balance the equation.



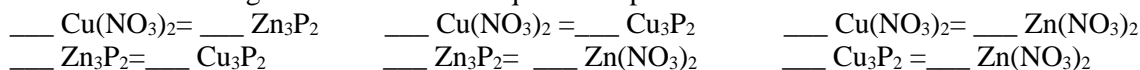
15. Give the following mole ratios for the equation in problem #14



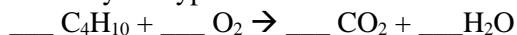
16. Identify the type of reaction and balance the equation.



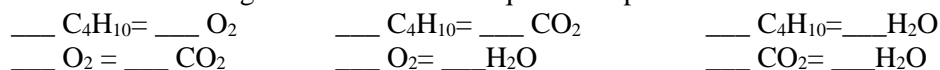
17. Give the following mole ratios for the equation in problem #16



18. Identify the type of reaction and balance the equation.



19. Give the following mole ratios for the equation in problem #18



20. Identify the type of reaction and balance the equation.



21. Give the following mole ratios for the equation in problem #20

