**Chemistry** **Name**
**Period** _____

# Writing Compound Formulas Worksheet

Write the correct formula for each of the following ionic compound names.

1. Copper (I) phosphate _____
2. Magnesium nitrite _____
3. Cobalt (II) hydroxide _____
4. Iron (III) chromate _____
5. Chromium (II) chloride _____
6. Zinc (II) fluoride _____
7. Mercury (II) acetate _____
8. Manganese (II) iodide _____
9. Copper (II) nitrate _____
10. Aluminum hydroxide _____
11. Iron (II) phosphate _____
12. Magnesium sulfite _____
13. Tin (II) bromide _____
14. Sodium bicarbonate _____
15. Chromium (III) oxalate _____
16. Iron (III) sulfate _____
17. Silver (I) dichromate _____
18. Lead (II) nitride _____
19. Strontium hypochlorite _____
20. Barium chlorate _____
21. Tin (IV) oxide _____
22. Nickel (II) cyanide _____
23. Aluminum sulfide _____
24. Lead (IV) chromate _____
25. Potassium permanganate _____
26. Calcium dihydrogen phosphate _____
27. Barium hydrogen sulfate _____
28. Lithium monohydrogen phosphate _____
29. Silver (I) bisulfite _____
30. Ammonium carbonate _____

Write the Compound formulas for the covalent molecules below:

31. Dinitrogen oxide _____
32. Carbon tetraiodide _____
33. Dioxygen heptachloride _____
34. Phosphorus tribromide _____
35. Tetraphosphorus trisulfide _____
36. Sulfur hexafluoride _____
37. Dinitrogen pentoxide _____
38. Tetrasulfur decanitride _____

Write the compound name for the following covalent molecules.

39. $CS_2$ _____
40. $PF_5$ _____
41. $CBr_4$ _____
42. $P_4S_{10}$ _____
43. $N_2S_5$ _____
44. $S_2F_{10}$ _____
45. $Cl_2O_7$ _____
46. $I_2O_5$ _____

**Write formulas for the following temary ionic compounds.**

1. sodium phosphate _____
2. magnesium nitrate _____
3. sodium hydroxide _____
4. potassium cyanide _____
5. ammonium bromide _____
6. potassium dichromate _____
7. cesium chlorate _____
8. iron (III) sulfate _____

**Write formulas for compounds using these pairs of ions.**

9. lithium, carbonate _____
10. nitride, tin (IV) _____
11. sulfide, potassium _____
12. hydroxide, chromium (III) _____
13. calcium, phosphide _____
14. chromate, ammonium _____

**Write formulas for the following binary compounds.**

15. nitrogen tribromide _____
16. dichlorine monoxide _____
17. dinitrogen tetrafluoride _____
18. xenon dichloride _____
19. lead (IV) sulfide _____
20. phosphorus pentafluoride _____
21. strontium bromide _____
22. aluminum nitrite _____
23. copper (II) iodide _____
24. hydrobromic acid _____
25. dihydrogen sulfide _____
26. dinitrogen difluoride _____

**Give the charge on the following ions:**

27. nitrate _____ 28. oxide _____ 29. phosphate _____ 30. nitride _____

**Name the following compounds:**

31. $CuCN$ _____
32. $SiO_2$ _____
33. $SeF_4$ _____
34. $HgI_2$ _____
35. $SO_3$ _____
36. $H_3PO_4$ _____
37. $SrBr_2$ _____
38. $NiCl_2$ _____
39. $K_2SO_4$ _____
40. $FeCO_3$ _____
41. $NH_4HCO_3$ _____
42. $NBr_3$ _____
43. $Cr(OH)_3$ _____
44. $NiI_2$ _____
45. $SnCl_4$ _____
46. $SF_6$ _____
47. $HCl$ _____
48. $AlF_3$ _____
49. $Cl_2O_7$ _____
50. $MnSO_4$ _____

**51. Name and write symbols for all seven diatomic molecules:**

a. _____
b. _____
c. _____
d, _____
e. _____
f. _____
g. _____